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CHAPTER 4 REVIEW Arrangement of Electrons in Atoms SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. d How many quantum numbers are used to describe the properties of electrons in atomic orbitals? (a) 1 (c) 3 (b) 2 (d) 4 2. a A spherical electron cloud surrounding an atomic nucleus would best represent (a) an s ...

4 Arrangement of Electrons in Atoms

CHAPTER 2 REVIEW Measurements and Calculations SECTION 1

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SHORT ANSWER Answer the following questions in the space provided. 1. Determine whether each of the following is an example of observation and data, a theory, a hypothesis, a control, or a model. observation and data a. A research team records the rainfall in inches

2 Measurements and Calculations

CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided.

1. Classify each of the following as a homogeneous or heterogeneous substance. homogeneous a. sugar homogeneous d. plastic wrap homogeneous b. iron filings heterogeneous e. cement sidewalk heterogeneous c. granola bar 2. For each type of investigation, select the most appropriate ...

1 Matter and Change - HUBBARD'S CHEMISTRY

CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter

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SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

3 Atoms: The Building Blocks of Matter

CHAPTER 5 REVIEW The Periodic Law SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. c When an electron is added to a neutral atom, energy is (a) always absorbed. (c) either absorbed or released. (b) always released. (d) neither absorbed nor released. 2. d The energy required to remove an electron from a neutral atom is the atom's

5 The Periodic Law

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 1 SHORT ANSWER Answer the following

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questions in the space provided. 1. c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom. (c) the charge on each Fe ion.

7 Chemical Formulas and Chemical Compounds

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CHAPTER 6 REVIEW Chemical Bonding SECTION 3 SHORT ANSWER Answer the following questions in the space provided.

1. a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. d In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b) ...

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6 Chemical Bonding

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attract electrons from another atom in the compound is called _____. 17. The energy required to remove one electron from an atom is called its _____. 18.

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